Synthetic Methods

Post-Synthetic Mannich Chemistry on Metal-Organic Frameworks: System-Specific Reactivity and Functionality-Triggered Dissolution


Abstract: The Mannich reaction of the zirconium MOF \( [\text{Zr}_6\text{O}_4(\text{OH})_4(\text{bdcd-NH}_2)] \) (UIO-66-NH\(_2\)) with paraformaldehyde and pyrazole, imidazole or 2-mercaptopimidazole led to post-synthetic modification (PSM) through C–N bond formation. The reaction with imidazole (Him) goes to completion whereas those with pyrazole (Hpzy) and 2-mercaptopimidazole (HimSH) give up to 41 and 36% conversion, respectively. The BET surface areas for the Mannich products are reduced from that of UIO-66-NH\(_2\) but the compounds show enhanced selectivity for adsorption of CO\(_2\) over N\(_2\) at 273 K. The thiol-containing MOF adsorb mercury(II) ions from aqueous solution, removing up to 99%. The Mannich reaction with pyrazole succeeds on \([\text{Zn}_2\text{O}(\text{bdcd-NH}_2)]\) (IRMOF-3), but a similar reaction on \([\text{Zn}_2(\text{bdcd-NH}_2)_2(\text{dabco})]\) (dabco = 1,4-diazabicyclo[2.2.2]octane) gave \([\text{Zn}_3(\text{bdcd-NH}_2)_2(\text{bdcd-NHCH}_2\text{pz})]_{13}\) (dabco) = 2-C,H\(_5\)5, whereas the reaction with imidazole gave the expected PSM product. Compound 5 forms via a dissolution-recrystallisation process that is triggered by the “free” pyrazolate nitrogen atom competing with dabco for coordination to the zinc(II) centre. In contrast, the “free” nitrogen atom on the imidazole is too far away to compete in this way. Mannich reactions on \([\text{In(OH)}(\text{bdcd-NH}_2)]\) (MIL-68(In)-NH\(_2\)) stop after the first step, and the product was identified as \([\text{In(OH)}(\text{bdcd-NH}_2\text{NHCH}_2\text{O})_3\text{bdc-N} = \text{CH}_2\text{bdc-N}]\), with addition of the heterocycle prevented by steric interactions.

Introduction

Metal-organic frameworks (MOFs)[1] are currently attracting considerable interest for their porosity properties, and applications as diverse as carbon capture,[2] catalysis,[3] drug delivery,[4] and chemical weapon detoxification.[5] Much of this attention arises from the wide diversity of MOF structures, with variation of both the metal centres and organic linkers providing an essentially limitless number of possible materials. Of specific interest for many applications is the potential for forming functionalised MOFs,[6] with particular functional groups appended to the pore walls. While such materials can sometimes be formed using a linker containing an appropriate substituent in the MOF synthesis, in practice many functional groups are intolerant to the synthetic conditions, or use of the functionalised linker in the synthesis gives rise to an unexpected product. Post-synthetic modification (PSM)[7] has emerged as a powerful tool for preparing such functionalised MOFs, and it is often the only way to place a particular substituent onto the pore walls of a MOF structure. A wide range of covalent post-synthetic modification reactions have been developed over recent years, including conversion of primary amines into amides,[8] isocyanates,[9] ureas,[10] azides,[11] β-amidoketones,[12] secondary amines[13] and diazonium salts,[14] aldehydes into hydrazones,[15] azides to triazoles,[16] bromides to nitriles,[17] as well as oxidation[18] and reduction[19] reactions. Despite this, there remains a need for new, versatile and synthetically-straightforward methods that allow different functional groups to be incorporated into MOFs, regardless of their metal centres and framework structure.

The Mannich reaction, first reported over 100 years ago,[20] involves the condensation of an amine with an aldehyde, normally formaldehyde, and a compound containing an active hydrogen.[21] Originally, this latter compound was an enolisable carbonyl such as an ester or a ketone, but development of the reaction has seen other nucleophiles such as nitroalkanes,[22] acetylenes[23] and electron-rich heterocycles, including pyrroles,[24] furans[25] and thiophenes,[26] being employed as alter-
nates to carbonyl compounds. In this paper, we explore the post-synthetic modification of the amino-functionalised metal-organic frameworks \([\text{Zr}_6\text{O}_4(\text{OH})_4(\text{bdc-NH}_2)_3]\) (Uio-66-NH\(_2\)), bdc-NH\(_2\) = 2-amino-1,4-benzene dicarboxylate,[31] \([\text{Zn}_2\text{O}(\text{bdc-NH}_2)_3]\) (IRMOF-3),[32] \([\text{Zn}_2(\text{bdc-NH}_2)_3(\text{dabco})]\) (DMOF-1-NH\(_2\), dabco = 1,4-diazabicyclo[2.2.2]octane),[32] and \([\text{In}(\text{OH})(\text{bdc-NH}_2)_3]\) (MIL-68(Ni), NH\(_2\)) using the Mannich reaction, employing pyrazole, imidazole and 2-mercaptoimidazole as the nucleophiles. The products from these transformations were anticipated to have nitrogen and/or sulfur groups projecting into the pores and available for selective gas adsorption or metal ion uptake. In all cases presented herein, the Mannich reaction was carried out in two steps to prevent the nucleophile from reacting with formaldehyde, and no catalyst was required.

Results and Discussion

Mannich reactions on \([\text{Zr}_6\text{O}_4(\text{OH})_4(\text{bdc-NH}_2)_3]\), Uio-66-NH\(_2\)

\([\text{Zr}_6\text{O}_4(\text{OH})_4(\text{bdc-NH}_2)_3]\), Uio-66-NH\(_2\) is an attractive PSM precursor due to the high chemical stability of the zirconium-dicarboxylate framework, its high crystallinity and relatively large pore windows (\(\approx 6 \text{ Å}\)) and the presence of the readily-functionalised amino groups.[32] Mannich reactions on Uio-66-NH\(_2\) were undertaken as shown in Scheme 1.

The first step involves the formation of methoxymethyl amine groups by the reaction with paraformaldehyde and MeOH at 50°C. These methoxymethyl amine groups were subsequently converted into the final product by reaction with pyrazole, imidazole or 2-mercaptoimidazole to give compounds 1–3, respectively. All reactions proceeded without the need for a Lewis acid catalyst, which has the additional advantage of eliminating the work-up associated with catalyst removal from the pores of the MOF and removes the possibility of pore blocking by the catalyst. The similarity between the PXRD patterns of Uio-66-NH\(_2\), and the PSM products 1–3 (Figure S2, S6 and S9) indicate that the original framework was maintained in all three cases.

The effectiveness of the PSM reactions in terms of the percentage conversion of amino groups into the Mannich products was gauged by \(^1\text{H}\) NMR spectroscopy. The \(^1\text{H}\) NMR spectra were obtained from MOF samples that were washed to remove unreacted reagents before digesting in NH\(_4\)F/D\(_2\)O with [D\(_2\)]DMSO. For the reaction with pyrazole (Hpz), the \(^1\text{H}\) NMR spectrum of 1 (Figure S3) shows a number of new signals in addition to those corresponding to the aromatic protons of the unmodified groups, present as D\(_2\)bdc-NH\(_2\) (\(\delta = 7.56\text{d}, 7.12\text{s} \) and 7.05 ppm). The aromatic protons of D\(_2\)bdc-NHCH\(_2\)pz were observed at \(\delta = 7.62\text{d}, 7.25\text{s} \) and 7.08d ppm, overlapping with the signals from D\(_2\)bdc-NH\(_2\) and others attributed to minor (<10%) by-products. The presence of the pyrazole ring on the digested framework of 1 was confirmed by the signals at \(\delta = 7.57\) and 6.28 ppm. Attempts to remove the by-products by thorough washing with a variety of solvents were unsuccessful, suggesting that these compounds are also derived from PSM reactions, with a double-Mannich product the most likely. Formylated by-products can be present in Uio-66-NH\(_2\), deriving from reaction with DMF during the MOF synthesis.[33] NMR analysis on digested samples of Uio-66-NH\(_2\) showed no evidence for formylation, suggesting this is not the origin of the by-products present in 1.

By comparison of the integrals for the signals at \(\delta = 7.13\) and 6.28 ppm, the percentage conversion from –NH\(_2\) into –NHCH\(_2\)pz groups was estimated to be 41%. Ignoring the minor by-products, this gives the formula for 1 as \([\text{Zr}_6\text{O}_4(\text{OH})(\text{bdc-NH}_2)_3(\text{bdc-NHCH}_2\text{pz})_2]\). Attempts to increase the degree of conversion by carrying out the reaction at a higher temperature or for a longer time period were unsuccessful, though it should be noted that higher conversion to the methoxymethyl amine in the first step might not be observable in the \(^1\text{H}\) NMR spectra of the digested product, given the likely reversion of any D\(_2\)bdc-NHCH\(_2\)OMe to D\(_2\)bdc-NH\(_2\) under the acidic digestion conditions.

The Mannich reaction of Uio-66-NH\(_2\) with imidazole (Him) as the nucleophile was more successful than that with pyrazole, with the amino groups fully converted into –NHCH\(_2\)im groups. This was confirmed by the disappearance of the signals which correspond to the aromatic protons of the starting MOF, Uio-66-NH\(_2\), in the \(^1\text{H}\) NMR spectrum of the digested product. Instead, new signals at \(\delta = 7.56\text{d}, 7.14\text{s} \) and 7.07d ppm were observed (Figure 1), corresponding to the protons from the benzene ring of D\(_2\)bdc-NHCH\(_2\)im. Furthermore, the presence of the imidazole ring can be confirmed by the presence of two singlets in the aromatic region (\(\delta = 7.75\) and 7.03 ppm). The signal at \(\delta = 7.03\) ppm corresponds to two chemically similar...
but non-identical protons from the imidazole ring, and this overlaps with the doublet from one of thearyl protons, where-
as the singlet at \( \delta = 7.75 \) ppm arises from the remaining proton peak of the imidazole ring. The signal attributed to the
methylene protons can be seen at \( \delta = 6.56 \) ppm, close to the broad HDO peak resulting from the digestion solvent. The
chemical formula of this PSM product is \([\text{Zr}_6\text{O}_5(\text{OH})_2(\text{bdc-}\text{NHCH}_2\text{imSH})_2]_2\).

In contrast to the complete conversion observed for 2, the comparable Mannich reaction with 2-mercaptoimidazole
(HimSH) as the nucleophile gave only partial conversion. The
\(^1\)H NMR spectrum (Figure S10) of the digested product 3
shows the presence of new peaks in addition to the aromatic proton peaks which correspond to the starting MOF, UiO-66-NH\(_2\). The signals attributed to the aromatic protons of D\(_2\)bdc-
NHCH\(_2\)imSH are observed at \( \delta = 7.68 \) d, 7.26 s and 7.08 d ppm, respectively, although these peaks overlap with others from
minor by-products. The presence of new peaks at \( \delta = 6.98 \) and 6.76 ppm, from the imidazole ring, indicates that the 2-mer-
captoimidazole ring was successfully grafted onto the MOF
framework.

The percentage conversion from -NH\(_2\) into -NHCH\(_2\)imSH was
calculated as approximately 36% by comparing the integrals
for the signals at \( \delta = 7.16 \) and 6.76 ppm. Ignoring minor by-
products, this gives a formula for 3 of \([\text{Zr}_6\text{O}_5(\text{OH})_2(\text{bdc-NH}-\text{NHCH}_2\text{imSH})_{13}]_2\).

For 1–3, further evidence for successful PSM came from the
ESI mass spectra of the digested products. The negative ion
ESI mass spectra of digested 1 and 3 confirmed the presence of
the deprotonated anions of \( \text{H}_2\text{bdc-NHCH}_2\text{pyz} \) and \( \text{H}_2\text{bdc-NHCH}_2\text{imSH} \)
at \( m/z = 260.0664 \) (predicted \([M-H]^- = 260.0671\))
and \( m/z = 292.0400 \) (predicted \([M-H]^- = 292.0392\)), respective-
ly. In both cases a peak was also observed for \( \text{H}_2\text{bdc-NH}_2 \)
\( (m/z = 180.0308, \text{predicted} \ [M-H]^+ = 180.0297) \). Digested 2
gave better results in the positive ion rather than the negative ion
ESI mass spectrum, with the protonated cation of \( \text{H}_2\text{bdc-NHCH}_2\text{im} \) at
\( m/z = 262.0824 \) (predicted \([M+H]^+ = 262.0828\)).

The percentage conversions for the PSM reactions generat-
ing 1–3 are summarised in Table 1. The differences in degree
of conversion can be related to the nucleophile strength. Imi-
dazole is a stronger nucleophile than pyrazole due to its
higher basicity, and is therefore more susceptible to nucleo-
philic substitution with –NHCH\(_2\)OCH\(_3\) leading to a higher con-
version. The steric demands of the nucleophile also have some
influence on the extent of the reaction, with the lowest con-
version achieved in the case of 2-mercaptoimidazole, the larg-
est of the nucleophiles employed. This can be rationalised by
the more restricted diffusion of 2-mercaptoimidazole within
the pores of the MOF.

The thiol substituent in 3 was anticipated to be able to coor-
dinate to soft metal centres such as mercury(II). In order to
probe the effect of different -NHCH\(_2\)imSH loadings on Hg\(^{II}\)
uptake, a second thiol-containing MOF was prepared, using
the same conditions as for 3, but with the temperature for
the second step reduced from 80 to 50 °C. It was anticipated that
the lower temperature during the second step would lead to a
lower conversion to the -NHCH\(_2\)imSH group.

The \(^1\)H NMR spectrum (Figure S13) of the digested product
formed under these conditions, 3a, showed the presence of the
modified group (-NHCH\(_2\)imSH), though present in a lower
relative concentration than in 3. The percentage conversion from
-NH\(_2\) into -NHCH\(_2\)imSH groups was estimated as 21%,
giving a formula for 3a of \([\text{Zr}_6\text{O}_5(\text{OH})_2(\text{bdc-NH}-\text{NHCH}_2\text{imSH})_{13}]_2\). This confirms that the reaction temperature
has a significant impact on the degree of modification, with a
lower temperature leading to lower conversion.

The TGA profiles of the PSM products 1–3 and 3a exhibit similar features to that for UiO-66-NH\(_2\) (Figure S14). There is an
initial mass loss (up to 110 °C) corresponding to removal of 1,4-
dioxane from the pores. A small, gradual mass loss, observed
in the range 110–470 °C, is attributed to the loss of residual sol-
vent in the pores and/or the dehydroxylation of the \( \text{Zr}_6\text{O}_5(\text{OH})_4 \)
nodes.\(^{[24]}\) The final mass loss, beginning at 470 °C, is due to the
decomposition of the framework. Based on the TGA profiles, 1
has 4.0, 2 has 3.0, 3 has 5.0, 3a has 5.5 and UiO-66-NH\(_2\) has
7.0 molecules of 1,4-dioxane per \( \text{Zr}_6\text{O}_5(\text{OH})_4 \) unit in the unacti-
vated MOFs. This shows that the amount of 1,4-dioxane in the
pores decreases as the degree of post-synthetic modification
increases. This is unsurprising, since the greater the degree of
conversion, the lower the residual space available to accom-
modate guest solvent molecules.

The BET surface areas of 1–3 and 3a were determined based
on their \( N_2 \) adsorption isotherms at 77 K (Figure 2). The com-
pounds were activated using the conventional activation tem-
perature for UiO-66 and its derivatives (120 °C for 12 h), and
the BET surface area for UiO-66-NH\(_2\) obtained in this work
(\( \delta_{\text{BET}} = 1041 \) m\(^2\)g\(^{-1}\)) is similar to previously reported values.\(^{[25]}\)
All PSM products exhibit type I isotherms, indicative of micro-

<table>
<thead>
<tr>
<th>Compound</th>
<th>Nucleophile</th>
<th>% Conversion</th>
</tr>
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<tbody>
<tr>
<td>1</td>
<td>pyrazole</td>
<td>41</td>
</tr>
<tr>
<td>2</td>
<td>imidazole</td>
<td>100</td>
</tr>
<tr>
<td>3</td>
<td>2-mercaptoimidazole</td>
<td>36</td>
</tr>
</tbody>
</table>

Figure 2. \( N_2 \) sorption isotherms for compounds 1–3 and 3a at 77 K, in compa-
rison to that for UiO-66-NH\(_2\).
porous materials, and have lower BET surface areas than UiO-
66-NH$_2$, with $S_{\text{BET}}$ values of 528 m$^2$g$^{-1}$ for 1, 290 m$^2$g$^{-1}$ for 2, 352 m$^2$g$^{-1}$ for 3 and 608 m$^2$g$^{-1}$ for 3a. BET surface areas are
governed by the degree of conversion and the size of the
modified groups. In general, the BET surface area reduces as
the percentage conversion increases and 2, with complete
conversion, has the lowest surface area. The presence of larger
pore groups in the pores also leads to lower BET surface
areas, with the value for 3 less than that for 1, despite 1 pos-
sessing a higher degree of modification.

The CO$_2$ adsorption isotherms of the PSM products were
measured at 273 K (Figure S15) to assess the influence of the
modified groups on the CO$_2$ uptake capacities. All PSM pro-
ducts show lower CO$_2$ uptake capacities than UiO-66-NH$_2$, attrib-
utable to the reduction in pore volume and the lower percent-
age of --NH$_2$ groups in the pores. Of the PSM products, 1
shows the highest CO$_2$ uptake which is probably due to the fa-
vourable interactions of CO$_2$ molecules with the nitrogen atom
in the pyrazole ring. Compound 2 shows a lower CO$_2$ uptake
than 1, despite having higher percentage of heterocycles in
the pores, which is consistent with the lower BET surface area,
ityl a consequence of the high degree of modification. Com-
pounds 3 and 3a show the lowest CO$_2$ uptake capacities at
1 bar and this may be due to pore blocking caused by higher
steric hindrance of the modified groups. Nonetheless, the pro-
portion of thiol groups in the pores has little impact on the
CO$_2$ uptake capacities, as evidenced by the relatively small dif-
fERENCE in CO$_2$ uptake between 3 and 3a. The modified MOFs
show enhanced CO$_2$/N$_2$ selectivity over UiO-66-NH$_2$, though
this is largely a consequence of their low N$_2$ uptake at 273 K.

The thiol-containing PSM products, [Zr$_2$O$_4$(OH)$_2$(bdc-
NH$_2$)$_2$,n(bdc-NHCH$_2$imNH)$_{15}$] 3 and [Zr$_2$O$_4$(OH)$_2$(bdc-
NH$_2$)$_2$,n(bdc-NHCH$_2$imSH)$_{15}$], 3a were also investigated for their
ability to remove mercury(II) from aqueous solutions. The Hg$^{II}$
uptake experiments were carried out by immersing the MOF
samples in an aqueous solution of HgCl$_2$ (100 ppm) and stir-
ringing the solution for 12 h at ambient temperature. The Hg$^{II}$-
treated MOFs were isolated by centrifugation and atomic emis-
sion spectroscopy (AES) was used to quantify the residual Hg$^{II}$
concentration in the supernatant.

Mercury uptake capacities were calculated using [Equa-
tion (1)] where $C_{\text{i}}$ and $C_{\text{f}}$ represent the initial and equilibration
Hg$^{II}$ concentrations, respectively. In addition to PSM products 3
and 3a, the Hg$^{II}$ uptake capacities of the unmodified MOFs,
UiO-66 and UiO-66-NH$_2$, were investigated for comparison,
with the results presented in Table 2.

\[
\text{Hg(II) uptake} = \left(\frac{C_{\text{i}} - C_{\text{f}}}{C_{\text{i}}} \times 100\right) \quad (1)
\]

The post-synthetic grafting of thiol groups in the pores of
UiO-66 proved to be beneficial for Hg$^{II}$ absorption, as the
uptake capacities were significantly increased for 3 and 3a
over the unmodified MOFs. Perhaps surprisingly, the highest
Hg$^{II}$ uptake was observed for 3a, despite 3 having a higher
loading of thiol groups in the pores. This reflects the lower po-
rosity of 3, which is likely to lead to some of the thiols being
unavailable to interact with the Hg$^{II}$ ions. The Hg$^{II}$ uptake in 3a
is comparable to that reported for the previously reported de-
ratuive UiO-66-(SH)$_2$ (34) which is one of the highest reported
for a MOF, demonstrating the potential of 3a for mercury re-
moval. PXRD (Figure S17) confirmed that 3a retains its crystal-
linity on treatment with HgCl$_2$ (aq).

Mannich reactions on [Zn$_4$O(bdc-NH$_2$)$_2$], IRMOF-3
IRMOF-3 contains large channels ($\approx$ 9.6 Å) and there is consid-
erable precedence for the post-synthetic modification of the
amino groups that protrude into its pores. As IRMOF-3 has a
low stability towards moisture and alcohol, toluene was se-
lected as the optimum solvent for the Mannich reaction.

To demonstrate the applicability of Mannich reaction on
IRMOF-3, the PSM reaction with pyrazole was carried out using
the reaction conditions outlined in Scheme 2.

The effectiveness of the PSM reaction was gauged by
$^1$H NMR spectroscopy on the DC/D$_2$O-digested product 4 (Fig-
ure S19). In addition to the signals corresponding to the aryl
protons of D$_2$bdc-NH$_2$, new features attributed to the aryl pro-
tons of the modified product were observed at $\delta = 7.89$d,
7.46$d$ and 7.20$d$ ppm. The successful incorporation of the
NHCH$_2$pyz groups could also be evidenced by the emergence
of new signals at $\delta = 7.85$d, 7.72$d$ and 6.25$d$ ppm, cor-
responding to the protons of the pyrazole ring. The peak attrib-
uted to the methylene protons was located at $\delta = 5.68$d. The
degree of conversion was calculated by comparing the in-
tegrals at $\delta = 7.46$d and 7.42$d$ ppm and found to be 75%, giving
the formula for 4 as [Zn$_4$O(bdc-NH$_2$)$_2$,n(bdc-NHCH$_2$pyz)$_2$]$_{23}$

The negative ion ESI mass spectrum of the digested product 4
confirmed the presence of the deprotonated anions of H$_2$bdc-
NHCH$_2$pyz and H$_2$bdc-NH$_2$ at $m/z = 260.0669$ (predicted
[M$-\text{H}^+]=260.0671$) and $m/z = 180.0308$ (predicted $[M$-
$H^+]=180.0308$).

<p>| Table 2. The Hg$^{II}$ uptake capacities of UiO-66, UiO-66-NH$_2$, and compounds 3 and 3a. |
|-------------------|-------------------|-------------------|-------------------|</p>
<table>
<thead>
<tr>
<th>Compound</th>
<th>$C_{\text{i}}$, ppm</th>
<th>$C_{\text{f}}$, ppm</th>
<th>Hg$^{II}$ uptake [%]</th>
</tr>
</thead>
<tbody>
<tr>
<td>UiO-66</td>
<td>100</td>
<td>89</td>
<td>11</td>
</tr>
<tr>
<td>UiO-66-NH$_2$</td>
<td>100</td>
<td>77</td>
<td>23</td>
</tr>
<tr>
<td>3</td>
<td>100</td>
<td>50</td>
<td>50</td>
</tr>
<tr>
<td>3a</td>
<td>100</td>
<td>1</td>
<td>99</td>
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</table>

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Mannich reactions on \([\text{Zn}_2\text{(bdc-NHCH}_2\text{dabco)}]\), DMOF-1-NH\(_2\)

\([\text{Zn}_2\text{(bdc-NHCH}_2\text{dabco)}]\), DMOF-1-NH\(_2\) is a flexible MOF which consists of \(\text{Zn}_2\)\((\text{dicarboxylate})\) sheets that are linked by dabco pillars into a three-dimensional network.\(^{29}\) MOFs in this series are able to undergo transitions from narrow rhomboidal pores to open, square pores, and this can be influenced by solvent or substituent.\(^{30}\) Toluene was selected as a solvent for post-SBU, and the gross structure of the framework SBU exists as sheets that are linked by dabco \((\text{Figure S22})\), revealing a significant structural difference between this peak and that for HDO, present from the digestion mixture. Comparing the integrals of the protons at \(\delta = 7.48–7.49\) ppm and \(\delta = 6.25\) ppm, the percentage conversion of amino into -NHCH\(_2\)pyz groups was calculated to be 56%.

The negative ion ESI mass spectrum of the digested product S shows the presence of aromatic protons attributed to \(\text{D}_6\text{bdc-NH}_2\) \((\delta = 7.82, 7.48\) and 7.13 ppm) and \(\text{D}_4\text{bdc-NHCH}_2\text{pyz} \((\delta = 7.89, 7.49\) and 7.20 ppm). The peaks at \(\delta = 7.13\) and 7.20 ppm overlap with the signals from the aryl protons of residual toluene solvent. The protons of the pyrazole ring are located at \(\delta = 7.85\) 7.72, and 6.25 ppm. The peak attributed to the \(\alpha\)-CH\(_2\) protons is observed at \(\delta = 5.68\) ppm although there is some overlap between this peak and that for HDO, present from the digestion mixture.

The negative ion ESI mass spectrum of the digested product S confirms the presence of the deprotonated anions of H\(_2\text{bdc-NHCH}_2\text{pyz}\) and H\(_4\text{bdc-NH}_2\) at \(m/z = 260.0662\) (predicted \([\text{M}–\text{H}] = 260.0671\) ) and \(m/z = 180.0364\) (predicted \([\text{M}–\text{H}] = 180.0375\) ), respectively. The disappearance of -NH\(_2\) stretching bands \((3287\) and 3457 cm\(^{-1}\)) of DMOF-1-NH\(_2\) in the FTIR spectrum of S \((\text{Figure S24})\) indicates the successful conversion of primary to secondary amine.

The XRD pattern of 5 is completely different to that of DMOF-1-NH\(_2\) \((\text{Figure S22})\), revealing a significant structural difference between the two materials. Indeed, the XRD pattern of 5 does not match any of the XRD patterns reported in the literature for DMOF-1 type materials. Inspection of 5 under an optical microscope revealed the presence of small colourless crystals and the absence of brown block crystals, characteristic of DMOF-1-NH\(_2\) and its derivatives. This observation suggests that DMOF-1-NH\(_2\) has undergone a complete structural change upon reaction.

The crystal structure of 5 was successfully elucidated by single crystal X-ray crystallography and is shown in Figure 3.
pared by Li and co-workers,\textsuperscript{39} contains zinc(II) metal centres which exhibit the same coordination geometries as those in 5.

In order to investigate the cause of the structural transformation from DMOF-1-NH$_2$ into 5, a series of control studies were carried out. No structural change was observed when DMOF-1-NH$_2$ crystals were heated in toluene, or when the crystals were treated separately with paraformaldehyde, MeOH or pyrazole (Figure S27). This suggested that the formation of the methoxymethyl amine intermediate DMOF-1-NHCH$_2$OCH$_3$ in the first step was unproblematic, but that the structural transformation occurred in the second step of the Mannich reaction. In order to confirm this, the reaction of DMOF-1-NHCH$_2$OCH$_3$ with pyrazole was monitored under an optical microscope equipped with a camera. The reaction conditions were modified in order to be able to view the reaction in this way. In particular, DMOF-1-NHCH$_2$OCH$_3$ crystals were dispersed on a microscope slide containing a solution of pyrazole in toluene at room temperature. After five minutes, the crystals began to dissolve, with complete dissolution observed after 40 minutes. A new phase, corresponding to the crystals of 5, was first observed after approximately twenty minutes (Figure 4), confirming that 5 is produced in a dissolution-re-precipitation process.

Although it is not possible to provide a definitive mechanism for the dissociation of DMOF-1-NHCH$_2$OCH$_3$, a proposed reaction mechanism that leads to the dissociation of the SBUs is shown in Figure 5.

After the first step of the Mannich reaction, the methoxymethyl amine species is localised in close proximity to the bridging dabco ligands. Upon addition of pyrazole, a facile re- action displacing methanol can occur to yield the -NHCH$_2$pyz group, which is aligned in such a way as to compete in an intramolecular manner with dabco for coordination to the Zn$^+$ metal centre. Displacement of dabco would break the three-di- mensional network of the DMOF-1 framework, leading to rapid delamination, and ultimately triggering framework dissolution. Notably, in the crystal structure of 5, the -NHCH$_2$pyz group is directed away from the dabco ligand (Figure S43), so is unable to compete with it for coordination. Moreover, a diaza-[18]- crown-6 ligand functionalised with pendant pyrazole groups using a Mannich reaction also exhibited fragmentation behaviour in the presence of transition metals,\textsuperscript{40} leading further credence to this hypothesis.

It should be noted that PXRD patterns for bulk samples of 5 show the presence of more than one phase, so the degree of occupancy of the pores by toluene in the crystal structure is an estimate, and although the ratio of linkers in the $^1$H NMR spectra are consistent between samples, this too may have been different in the crystal analysed crystallographically.

The Mannich reaction of DMOF-1-NH$_2$ with imidazole as the nucleophile was carried out using the same conditions as with pyrazole (Scheme 2). The $^1$H NMR spectrum of the digested product 6 (Figure S30) shows aromatic protons from D$_2$bdc-NH$_2$, and from the integrals the percentage conversion of amino into -NHCH$_2$im groups was calculated to be 65%, giving a formula for 6 as [Zn$_2$((bdc-NH)$_2$)$(\text{bdc-NHCH}_2$im)$_2$](dabco)]. The negative ion ESI mass spectrum of the acid-digested product 6 confirmed the presence of the deprotonated anions of H$_2$bdc-NHCH$_2$im and H$_2$bdc-NH$_2$ at $m/z = 260.0661$ (predicted [M$-$H$^+$] = 260.0671) and $m/z = 180.0339$ (predicted [M$-$H$^+$] = 180.0297).

The PXRD pattern of 6 and the starting MOF, DMOF-1-NH$_2$, closely match one another (Figure S29), demonstrated that PSM does not affect the gross structure or the crystallinity of the product. Furthermore, visual inspection of 6 confirmed the presence of only brown block crystals and the absence of new phases. Attempts to analyse 6 crystallographically were hampered by crystal twinning. Nonetheless, a screening experiment suggested that there were similarities in the unit cell parameters of 6 ($a = 15.2955(17)$ Å, $b = 15.2860(15)$ Å, $c = 19.207(2)$ Å) and those of DMOF-1 ($a = 15.063(2)$ Å, $c = 19.247(5)$ Å).

Based on these results, it is clear that framework dissolution does not occur when imidazole is used as a nucleophile. It is believed that substituting pyrazole by imidazole prevents the dissociation of DMOF-1-NHCH$_2$OCH$_3$, by eliminating the possibility of coordinative competition with dabco. The "free" nitrogen atom in imidazole is positioned beyond the coordination sphere of the zinc(II) centre and, as a consequence, the process shown in Figure 5 is unable to occur.

The Mannich reaction of DMOF-1-NH$_2$ with 2-mercaptoimidazole as the nucleophile was attempted using the same conditions as in the reaction with imidazole. However, the $^1$H NMR spectrum of the digested solid showed only signals corresponding to the aryl protons of DMOF-1-NH$_2$ (Figure S37), indicating that the inclusion of 2-mercaptoimidazole onto this MOF framework was unsuccessful. The PXRD pattern (Figure S36) is similar to that for DMOF-1-NH$_2$, implying that the framework was retained throughout the experiment. The unsuccessful graft-
Mannich reactions on [In(OH)(bdc-NH₂)]₄, MIL-68(In)-NH₂

[In(OH)(bdc-NH₂)]₄ MIL-68(In)-NH₂, is a three-dimensional MOF that is constructed from chains of InO₆(OH)₃ octahedral units that are linked together by bdc-NH₂ ligands to form triangular (≈ 6 Å) and hexagonal (≈ 16 Å) one-dimensional channels. In MIL-68(In)-NH₂, the amino groups are oriented towards the InO₆(OH)₃ octahedral chains rather than projecting into the pores. However, this has not prevented successful tandem post-synthetic modifications involving formation of the azide and subsequent click reactions from being carried out,[31] so presumably some flexibility is possible to accommodate the bulkier, modified groups.

MIL-68(In)-NH₂ was prepared using an analogous synthesis to that for MIL-68(In), originally reported by Loiseau and co-workers.[31] In a typical PSM procedure, MIL-68(In)-NH₂ crystals were treated with paraformaldehyde and MeOH at 50 ℃ for 24 h. In this reaction, MeOH was used as a reactant as well as a solvent, as MIL-68(In) is stable towards alcohols, thus eliminating the need to use a different solvent. The intermediate product was then washed with 1,4-dioxane and treated with pyrazole at 80 ℃ for 24 h, before quenching the reaction by washing the sample with fresh 1,4-dioxane.

The ¹H NMR spectrum of the digested PSM product 7 (Figure S39) was obtained by digesting the MOF in a basic aqueous solution (NaOD/D₂O). In addition to the signals corresponding to the aromatic protons of Dbdc-NH₃⁺, two new sets of signals were observed in the downfield region of the spectrum. However, the absence of peaks attributed to the protons of the pyrazole ring indicated that the PSM reaction did not afford the expected pyrazole-containing product. The signals at δ = 7.33d, 7.36s and 7.15d ppm are believed to be due to the aryl protons from the intermediate MOF, MIL-68(In)-NHCH₃OH, observed as Dbdc-NHCH₃OH⁺ in the NMR spectrum. The peaks attributed to the methylene protons and methyl terminus of Dbdc-NHCH₃CH₃⁺ are located at δ = 4.87 and 4.76 ppm, respectively, although these are partly obscured by the peak from HDO, present from the digestion solvent. The other signals, at δ = 7.68d, 7.43s and 7.18d ppm, are believed to be from the imine Dbdc-N = CH⁺, with mass spectrometry providing support for this (vide infra).

In order to confirm that the observed products do not require the presence of pyrazole, the reaction mixture was analysed prior to its addition. The first step of the Mannich reaction is depicted in Scheme 3, broken down into two stages. As anticipated, the ¹H NMR spectrum of the digested product (Figure S40) illustrates a high similarity with that for 7, with only small differences in the relative proportions of the two products.

This finding validates the hypothesis that 7 contains unreacted -NH₂ groups, as well as imine and methoxymethyl amine species. The presence of the imine could be due to incomplete reaction with methanol or, alternatively, from the partial hydrolysis of Dbdc-NHCH₃OH in the digestion medium. Given that Dbdc-NHCH₃OH appears to be stable under the digestion conditions, the most reasonable formulation for 7 includes both substituents, and can be represented by the formula [In(OH)(bdc-NH₂)₄]/(bdc-NH₂CH₃)[bdc-N = CH₂]₂.

The negative ion ESI mass spectrum of the base-digested product 7 (Figure S41) confirms the presence of the deprotonated anions of H₂bdc-NHCH₃OH, ([M–H]⁻ = 224.0560, predicted 224.0559), H₂bdc-N = CH₃ ([M–H]⁻ = 192.0310, predicted 192.0297) and H₂bdc-NH₂ ([M–H]⁻ = 180.0315, predicted 180.0297), and provides good evidence for the identity of the tag groups in the products. The PXRD pattern of 7 is similar to that of MIL-68(In)-NH₂ (Figure S38), indicating that framework integrity is maintained and the PSM reaction did not alter the crystallinity of the product.

Crystals were grown from dioxane, and the crystal structure of 7·0.8 dioxane was successfully elucidated by single crystal X-ray crystallography. The asymmetric unit (Figure S46) consists of two indium(III) centres with site occupancies of 0.5 and 0.25 for In1 and In2, respectively, one half and one quarter of a dicarboxylate ligand and two OH ligands (based on O1 and O5) with combined site occupancies of 0.75. Finally, there was evidence for some diffuse solvent present in the framework which was modelled as four-fifths of a dioxane molecule per indium centre based on TGA evidence.

As can be seen in Figure 6a,b, the overall framework topol-ogy has not changed significantly during the reaction from that of MIL-68(In)-NH₂, in agreement with the PXRD data. Although there is some evidence for the nitrogen atoms of the tag groups, disordered over several positions, further evidence for the nature of the substituents was unavailable.

The most notable insight from the crystal structure of 7 is the short distance between the nitrogen atoms on neighbouring benzene rings (Figure 6c). Although these atoms have only partial occupancy, this proximity illustrates the difficulties involved in placing a large substituent on one of these atoms. This provides justification for the argument that the second step of the Mannich reaction is disfavoured in this case on steric grounds.

Conclusion

The results presented herein demonstrate a previously unre-
ported post-synthetic modification process on MOFs, whereby catalyst-free Mannich reactions were used to convert the pri-
and MIL-78 into a range of azole-functionalised MOFs with conversion of up to 100%. It did not prove possible to prepare the azole-functionalised acids using similar procedures, and the instability of N–CH₂–N linkages to hydrolysis is well established. [...] Hence post-synthetic modification provides the only method to prepare these functionalised MOFs.

With regards to the PSM reactions on UiO-66-NH₂, the degree of conversion from -NH₂ into -NHR (R = CH₂-pyz, CH₂-im and CH₂-imSH) depends on the strength and size of the nucleophiles. Complete conversion was achieved with the strongest nucleophile (imidazole) whereas a lower conversion (41%) was obtained with the isosteric weaker nucleophile, pyrazole. The use of a larger nucleophile, 2-mercaptoimidazole, led to the lowest conversion (36%) and this is most likely due to the restricted diffusion of the nucleophile within the pores of UiO-66-NH₂. The modified MOFs have lower BET surface areas than UiO-66-NH₂, but show enhanced selectivity for CO₂ over N₂. In addition, the thiol-containing products show excellent uptake of mercury(II) from aqueous solutions.

With regard to the PSM reaction on IRMOF-3, 75% conversion of -NH₂ into -NHCH₂-pyz was achieved whilst using pyrazole as a nucleophile. However, the successful PSM reaction comes at a cost of decreased product crystallinity as evidenced by the broadening of peaks and reduction in peak intensities in the PXRD pattern of the PSM product.

The Mannich reaction on DMOF-1-NH₂ using pyrazole as the nucleophile, unexpectedly afforded [Zn₉(bdc-NH₂)₁₂(bdc-NHCH₂-pyz)₁₆(dabco)]·2C₃H₆, which was characterised by single crystal X-ray crystallography, ¹H NMR spectroscopy and TGA analyses. The framework transformation occurs when the intermediate MOF, DMOF-1-NHCH₂OCH₃ dissolves in the presence of pyrazole and re-precipitates. In contrast, the Mannich reaction of DMOF-1-NH₂ with imidazole afforded a product, 6, bearing the same gross structure as DMOF-1-NH₂, showing that substituting pyrazole for imidazole prevents the dissolution of DMOF-1-NHCH₂OCH₃. This difference in reactivity has been rationalised on the basis of a functionality-dependent dissolution process, in which the “free” nitrogen atom on pyrazole is in a position to compete with the dabco ligand for coordination to zinc, whereas the equivalent atom on imidazole is too far away to coordinate.

Subjecting MIL-68(In)-NH₂ to a similar PSM reaction with pyrazole, gave a modified product 7 that did not contain the heterocycle. The first step of the Mannich reaction proceeded, but the methoxymethyl amine intermediate did not react with pyrazole in the expected manner. The X-ray crystal structure of 7 suggests that this is a consequence of the location and orientation of these groups which are inaccessible to the pyrazole molecules, thus preventing the second step in the Mannich reaction from occurring.

This work has demonstrated that the post-synthetic Mannich reaction represents a versatile route to introducing complex functionalities into a range of metal-organic frameworks, and we are currently working to further develop the breadth of this approach.

**Experimental Section**

Full experimental details are presented in the electronic supplementary information. As an example, the reaction of UiO-66-NH₂ with formaldehyde and imidazole is presented here. UiO-66-NH₂ (117 mg, 0.4 mmol eq. of NH₂) and paraformaldehyde (24 mg, 0.8 mmol, 2 equiv.) were added into a glass vial containing methanol (5 mL). The vial was placed in an oven and heated at 50 °C for 24 h. The powder was then washed with methanol (three times) via centrifugation to remove any residual paraformaldehyde in the pores or on the solid surfaces. The powder was subsequently treated with imidazole (54 mg, 0.8 mmol, 2 equiv.) in 1,4-dioxane at 80 °C for 24 h before quenching the reaction by rinsing the sample.
with fresh 1,4-dioxane. The product was soaked in 1,4-dioxane for 3 days, replacing the solvent with fresh solvent every 24 h, before isolation by centrifugation. Prior to characterisation, samples were left to dry in air for 2 h to obtain free-flowing powders. Full details of the X-ray crystal structures of 5 and 7-0.8 dioxane are given in the Supplementary Information. The structures have also been deposited with the Cambridge Structural Database (CCDC 1824632:3 contains the supplementary crystallographic data for this paper. These data are provided free of charge by The Cambridge Crystallographic Data Centre).

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Conflict of interest
The authors declare no conflict of interest.

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